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Ka Kb Kw pH pOH
pKa pKb H+ OH-
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Equivalence Point,

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strong acids and 6
strong bases

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~~Acids + Bases~~

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~~What the Heck is~~

~~an Acid or Base?~~

~~Organic Chemistry~~

~~Chapter 16 - Acid-~~

~~Base Equilibria:~~

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~~115: Acid-Base~~

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main idea. acids
and bases change
the color of
compound~~

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-Lewis acid-base reaction is the formation of one or more covalent bonds between an electron-pair donor and an electron pair acceptor, although the 3 acid-base reactions differ, many compounds can be categorized as

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15: Acids and

Bases 1. The

hydronium ion and
the hydroxide ion,
in that order, are:

A) H_3O^+ , OH^-

B) OH^- , H_3O^+ C)

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OH-, H + D) H 3 O
+, OH-E) H 3 O-
OH-Ans: D

Category: Easy
Section: 15.1 2.

Which of the
following does not
fit the definition of
a Brønsted Acid?

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and Bases. strong
acid. strong base.
weak acid. weak
base. an acid that
ionizes completely
in solvent. a base
that ionizes
completely in a
solvent. an acid
that releases few
hydrogen ions in
aqueous solution. a
base that releases

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few hydroxide ions
in aqueous
solution.

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Answers~~

ACIDS AND BASES
453 SECTION 15-1
OBJECTIVES List
five general
properties of
aqueous acids and
bases. Name

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Common binary acids and oxyacids, given their chemical formulas.

List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius's theory of ionization.

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and Bases Acids

and Bases

Arrhenius

Definitions: acids -

compounds that

produce an

increase in $[H^+]$

when dissolved in

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water bases -
compounds that
produce an
increase in $[\text{OH}^-]$
when dissolved in
water Lewis
Definitions: acids -
electron pair
acceptors bases -
electron pair
donors Brønsted-
Lowry Definitions:
acids - H^+ donors

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Thus, H_2O and

H_3O^+ , and H_2O

and OH^- are

conjugate acid-

base pairs, but

H_3O^+ and OH^- are

not conjugate acid-

base pair. A

Brønsted-Lowry

acid-base reaction

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involves a competition between two bases for a proton, in which the stronger base ends up being the most protonated at equilibrium. In the reaction: $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$,

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acid. strong base.
weak acid. weak
base. an acid that
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in solvent. a base
that ionizes
completely in a
solvent. an acid
that releases few

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hydrogen ions in
aqueous solution. a
base that releases
few hydroxide ions
in aqueous
solution.

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mpahydrauliczna.e
†~~

Chapter 15 Review
Acids Bases Label
them as acid, base,

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conjugate acid (ca)
and conjugate base
(cb) Answer: CH_3NH_2 (aq) is the
base (an amine), H_2O (l) acts as the
acid here, CH_3NH_3^+ (aq) is the ca,
 OH^- (aq) is the cb;
Remember that H^+
+ is just short hand
for H_3O^+ called
hydronium ion. H^+
doesn't really exist

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Answer Key~~

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acid. strong base.
weak acid. weak
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completely in a

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solvent. an acid
that releases few
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aqueous solution. a
base that releases
few hydroxide ions
in aqueous
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Bronsted Acids and
Bases; Acids -
Bases Answer

molecules that can
lose H^+ (proton
donors) making H_3O^+
in water ...

Since x is $= [H^+]$
find the $pH = -\log$
 $7.0746 \times 10^{-3} =$
2.15 (we need 2
decimal places in
the final answer)

For more ...

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SHORT ANSWER

Answer the
following questions
in the space
provided. 1. Name
the following
compounds as
acids: sulfuric acid

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a. H_2SO_4 sulfurous acid
b. H_2SO_3 hydrosulfuric acid
c. H_2S perchloric acid
d. HClO_4 hydrocyanic acid
e. hydrogen cyanide

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~~– David Brearley~~
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Answers points.
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Key
more than
additional will meet
the expense of
each success.
neighboring to, the
statement as well
as insight of this
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answers can be

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taken as
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cleanup
professionals, and
concerned citizens.
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46a6cd3c5446b